Electromotive force series

ELECTRODE REACTION	STANDARD POTENTIAL (V VS. SHE)
$O_3(g) + 2H^+ + 2e^- \rightarrow O_2(g) + H_2O$	2.070
$H_2O_2 + 2H^+ + 2e^- \rightarrow 2H_2O$	1.776
$Au^{3+} + 3e^- \rightarrow Au$	1.500
$Cl_2(g) + 2e^- \rightarrow 2Cl^-$	1.358
$Pt^{2+} + 2e^- \rightarrow Pt$	1.200
$O_2 + 4H^+ + 4e^- \rightarrow 2H_2O$	**1.229**
$Pd^{2+} + 2e^- \rightarrow Pd$	0.987
$H_2O_2 + 2e^- \rightarrow 2OH^-$	0.880
$Hg^{2+} + 2e^- \rightarrow Hg$	0.854
$Ag^+ + e^- \rightarrow Ag$	0.800
$Cu^+ + e^- \rightarrow Pd$	0.521
$O_2 + 2H_2O + 4e^- \rightarrow 4OH^-$	**0.401**
$Cu^{2+} + 2e^- \rightarrow Cu$	0.337
$2H^+ + 2e^- \rightarrow H_2$	**0.000**
$Fe^{3+} + 3e^- \rightarrow Fe$	-0.036
$Pb^{2+} + 2e^- \rightarrow Pb$	-0.126
$Sn^{2+} + 2e^- \rightarrow Sn$	-0.136
$Ni^{2+} + 2e^- \rightarrow Ni$	-0.250
$Co^{2+} + 2e^- \rightarrow Co$	-0.277
$In^{3+} + 3e^- \rightarrow In$	-0.342
$Cd^{2+} + 2e^- \rightarrow Cd$	-0.403
$Fe^{2+} + 2e^- \rightarrow Fe$	-0.440
$Ga^{3+} + 3e^- \rightarrow Ga$	-0.530
$Cr^{3+} + 3e^- \rightarrow Cr$	-0.740
$Zn^{2+} + 2e^- \rightarrow Zn$	-0.763
$Mn^{2+} + 2e^- \rightarrow Mn$	-1.180
$Zr^{4+} + 4e^- \rightarrow Zr$	-1.530
$Al^{3+} + 3e^- \rightarrow Al$	-1.660
$Mg^{2+} + 2e^- \rightarrow Mg$	-2.370
$Na^{2+} + 2e^- \rightarrow Na$	-2.710
$Ca^{2+} + 2e^- \rightarrow Ca$	-2.870
$K^{2+} + 2e^- \rightarrow K$	-2.930
$Li^{2+} + 2e^- \rightarrow Li$	-3.050

References

1. R. W. Revie, H. H. Uhlig, “Corrosion and corrosion control: an introduction to corrosion science and engineering”, Chapter 3, p.31 (2008). Buy book at Amazon or Kindle Edition.
2. P. R. Roberge, “Corrosion Engineering: Principles and Plractice”, Chapter 4, Table 4.1, McGraw-Hill Professional (2008). Buy book at Amazon or Kindle Edition.